



Further Studies on the Softening of Rigid PVC by Aqueous Solutions of Organic Solvents

Louise V. Parker and Thomas A. Ranney

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Abstract: In this study, small pieces of PVC well casing were exposed to relatively low activities (or relative solubilities) of TCE and methylene chloride for 12 months. PVC pieces were also exposed to aqueous solutions containing several organic chemicals that were either solvents or swelling agents of PVC. In addition, small pieces of PVC were exposed to aque-

ous solutions containing organic chemicals that were either PVC solvents or swelling agents and were totally miscible in water. These studies revealed that there is an interactive effect among these chemicals when dissolved in water. However, softening does not appear to occur in any solutions where the sum of the relative solubilities is less than 0.1.

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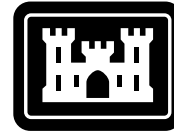
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PREFACE

This report was prepared by Louise V. Parker, Research Physical Scientist, Applied Research Division, Research and Engineering Directorate, U.S. Army Cold Regions Research and Engineering Laboratory (CRREL), Hanover, New Hampshire, and Thomas A. Ranney, Staff Scientist, Science and Technology Corporation, Hanover, New Hampshire. Funding for this work was provided by the U.S. Army Environmental Center (USAEC), Aberdeen Proving Ground, Maryland, Martin H. Stutz, Project Monitor.

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Further Studies on the Softening of Rigid PVC by Aqueous Solutions of Organic Solvents

LOUISE V. PARKER AND THOMAS A. RANNEY

INTRODUCTION

It has been known for some time that some neat (undiluted) organic solvents can dissolve or swell PVC well casings and screens. Products that can dissolve PVC are referred to as PVC solvents, and products that do not dissolve PVC but can cause it to swell are referred to as swelling agents of PVC. There has been concern that PVC casings and screens might be degraded by well water containing ppm and ppb levels of these chemicals. In the past few years, we have conducted a series of studies (Taylor and Parker 1990, Parker et al. 1992, Parker and Ranney 1994, 1995) to address this issue.

The solvent interaction parameter (χ) used in the Flory-Huggins equation* has been commonly used as a measure of the solvent or swelling power of a particular organic solvent (Berens 1985, Vonk 1985, 1986, Jenkins et al. 1986). Table 1 gives χ values for a number of organic chemicals that were derived experimentally[†] by Berens (1985) at 30°C and by Vonk (1985) at 20°C. PVC is reported to be softened only by solvents with a χ value less than 1 (Berens 1985) or 1.2 (Vonk 1985); the lower the number the greater the solvent or swelling power. Also, a neat solvent with a χ value less

Table 1. Literature values for the Flory-Huggins interaction parameter, χ .

<i>Organic chemical</i>	<i>From Berens (1985)</i>	<i>From Vonk (1985)</i>
acetone	0.61	—
aniline	—	0.87
benzene	0.83	0.83
2-chloroaniline	—	0.58
3-chloroaniline	—	0.61
chlorobenzene	—	0.66
chloroform	0.64	—
1,2-chlorotoluene	—	0.74
1,3-chlorotoluene	—	0.69
1,4-chlorotoluene	—	0.75
1,2-dichlorobenzene	—	0.63
1,3-dichlorobenzene	—	0.74
1,1-dichloroethane	< 0.68	—
1,2-dichloroethane	< 0.55	0.56
1,2-dichloropropane	0.68	0.65
di-isopropylketone	—	0.56
ethylbenzene	—	0.89
<i>n</i> -methylaniline	—	0.59
methylene chloride	< 0.53	0.55
methyl ethyl ketone	—	0.54
nitrobenzene	—	0.54
propylbenzene	—	1.91
1,1,2,2-tetrachloroethane	—	0.59
tetrachloroethylene	1.17	1.58
toluene	0.80	0.82
1,2,4-trichlorobenzene	—	1.52
1,1,1-trichloroethane	0.85	0.92
1,1,2-trichloroethane	< 0.56	0.58
trichloroethylene	0.88	0.90
1,3,5 trimethylbenzene	—	2.52
vinyl chloride	0.98	—
xylene	0.88	—
1,2-xylene	—	0.76
1,3-xylene	—	0.87
1,4-xylene	—	0.86

* The Flory-Huggins equation relates activity (a) to the volume fraction (V_1) of the organic chemical in the polymer as follows:

$$\ln(a) = \ln V_1 + (1 - V_1) + \chi(1 - V_1)^2$$

where χ is the Flory-Huggins interaction parameter.

[†] Berens (1985) and Vonk (1985) derived χ by exposing PVC to different activities of organic chemicals in vapor and solution.

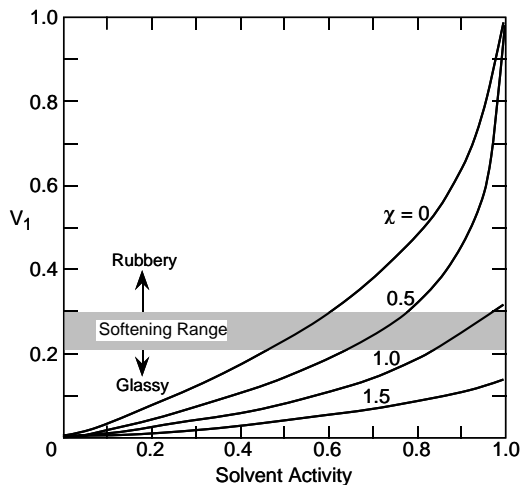


Figure 1. Flory-Huggins isotherms for various values of χ and approximate composition for softening of PVC at room temperature (from Berens 1985).

than 0.5 reportedly can completely dissolve PVC (Berens 1985). According to Vonk (1985), PVC is nearly inert to organic solvents if their χ value is greater than 3 (e.g., alcohols, aliphatic hydrocarbons, and organic acids).

According to Berens (1985), at ambient temperatures, rigid PVC can be softened only by PVC solvents or strong swelling agents at activities greater than 0.5. The less aggressive the solvent (i.e., the higher the χ value), the higher the activity required to cause softening. Figure 1 shows the isotherms Berens derived using the Flory-Huggins equation for various values of χ and the approximate softening range for PVC at 30°C. Berens predicted that for lower temperatures (such as those commonly encountered in groundwater), slightly higher activities would be needed to cause softening.

The activity of an organic chemical in aqueous solution can be approximated by dividing the concentration of the compound in solution (C_w) by its solubility in water (C_{ws}) (Berens 1985, Vonk 1986, Holsen 1988). However, the true relationship between activity (a), C_w , and C_{ws} is:

$$a = \gamma \frac{C_w}{C_{ws}}$$

where γ is the activity coefficient (Holsen 1988).

According to Holsen (1988), γ is equal to unity when there are no interacting forces between the molecules; these forces become important at high but still unsaturated conditions. At lower concentrations, γ has been considered equal to unity

partially because of the lack of experimental data. Therefore, at high concentrations or when there are interacting forces between the molecules, estimates of activity using solubility may be inaccurate.

In this paper, we will use the term “relative solubility” to express C_w/C_{ws} and thus estimate activity. However, we should point out that there also can be significant imprecision with this term. While it is known that aqueous solubilities vary with temperature, we have also found that for some organic compounds, the reported solubility can vary 1 or 2 orders of magnitude at the same temperature. Montgomery and Welkom (1990) have compiled aqueous solubility data from a number of sources. As an example of the variability we have observed, two of the values they report for the solubility of 1,1-dichloroethylene at 25°C are 273 mg/L and 5,000 mg/L.

Berens (1985) concluded that if aqueous solutions of solvents are going to cause softening of rigid PVC, the pure (or neat) solvent must be able to soften or dissolve PVC and must exist in solution at a concentration that exceeds one-half its solubility in water (activity > 0.5). However, one criticism of Berens’ work is that the various test (activities) solutions he used in his experiments were prepared by dissolving the organic solvents in polyethylene glycol and thus he never actually tested aqueous solutions of these organic solvents. Since one would expect that the partitioning of these organic solutes between the dissolved state and the polymer would differ between glycol and water, studies that use aqueous solutions of PVC solvents should be conducted to confirm Berens’ predictions. While Vonk did use aqueous solutions of organics in his studies, we were unable to find a thorough discussion of his methods that was available in English. Therefore, our laboratory conducted a series of studies (Parker et al. 1992, Parker and Ranney 1994, 1995) to confirm Berens’ predictions for PVC solvents and swelling agents in aqueous solutions. Following is a summary of our findings to date.

Methylene chloride studies

The first test solutions that were tested contained methylene chloride. Vonk (1985) reported that the χ value of methylene chloride is 0.55, while Berens (1985) did not give a value for χ but reported that it is less than 0.53. We (Ranney and Parker 1995) have observed pieces of PVC well casing dissolve when placed in neat methylene chloride. Thus, if Berens (1985) is correct in say-

ing that any organic chemical with a χ value less than 0.5 can dissolve PVC, the true χ value for methylene chloride remains undefined but is less than 0.5.

In a previous study (Parker and Ranney 1994), we found that after one week of exposure, PVC pieces exposed to aqueous solutions of methylene chloride with a relative solubility of 0.6 had hardness readings that were significantly lower than the control samples, and these samples were rubbery in that they could easily be bent by hand. (In all our studies, we use a Barcol Impressor to determine hardness.) Samples exposed to solutions with a relative solubility of 0.8 were also similarly affected, with significant changes in the hardness readings after only four hours. While these findings agree with Berens' (1985) predictions for PVC exposed to aqueous solutions with an activity of 0.8, Berens would not predict that PVC exposed to solutions with an activity, or relative solubility, of 0.6 would become rubbery unless the true value of χ was very close to 0. We were not able to find in Berens' paper how he determined softening.

In a subsequent study (Parker and Ranney 1994), we were able to determine that there were significant changes in the hardness readings of test pieces that were exposed to aqueous solutions of methylene chloride with relative solubilities of 0.4, 0.2, and 0.1 for 20 weeks. However, none of these samples were actually rubbery or pliable. We concluded that it appeared that softening had begun to occur at much lower relative solubilities, or activities (0.1 to 0.4), than had been previously reported or predicted by Berens (1985). Because we conducted our own solubility tests to confirm the aqueous solubility of methylene chloride, we do not feel that our test results can be explained away by assuming the solubility value we used is incorrect.

According to Berens (1985), diffusion of organic permeants is slowest (Fickian) at activities less than 0.25. However, Vonk (1985, 1986) felt that diffusion would be Fickian for the poorer PVC swelling agents (e.g., alkylated aromatics) at activities less than 0.25, but for the solvents and the better swelling agents (i.e., anilines, chlorinated hydrocarbons, ketones, and nitrobenzenes) the activity should be less than 0.1. If this is the case, we would not expect to observe any changes in the hardness readings of samples exposed to solutions of methylene chloride (a solvent) with relative solubilities less than 0.1. The results from this study tend to support this hypothesis since

the hardness readings of the samples that were exposed to solutions with a relative solubility of 0.05 were not significantly different from the controls.

We proposed that only partial softening occurs at the lower relative solubilities (0.1 to 0.4) and that these samples will never actually become rubbery or pliable. Examination of Berens' isotherms (Fig. 1) indicates that some samples do become partially softened but never actually become rubbery. We concluded that a longer-term study was necessary to resolve whether softening progresses further with time, and the lowest relative solubility at which softening occurs.

Trichloroethylene studies

We conducted another series of studies using trichloroethylene (TCE). TCE is a common environmental pollutant and is a relatively good swelling agent of PVC. There is close agreement between the χ value of Berens (1985), 0.88, and that of Vonk (1985), 0.90. When we (App. A, Ranney and Parker 1995) subjected small pieces of PVC well casing to neat (undiluted) TCE, we found that the test pieces showed signs of swelling and softening after only one day. However, TCE did not dissolve the PVC pieces. Thus, our experimental work appears to confirm that TCE is a good swelling agent of PVC but that it is not a solvent.

We conducted two studies where pieces of PVC were exposed to aqueous solutions of TCE with a range of relative solubilities (0.2 to 1.0) (Parker and Ranney 1994). Samples that were exposed to solutions with relative solubilities of 0.8 and 1.0 had Barcol readings that were significantly lower than the controls. The samples that were exposed to the solutions with a relative solubility of 0.8 never became pliable or rubbery, while those exposed to the 1.0 relative solubility solutions were pliable by the end of the study (77 days). Generally, the results for TCE agree with what would be predicted using Berens' isotherms (Fig. 1): PVC would be softened by TCE when the activity, or relative solubility, is 0.8 or greater and would become rubbery if the relative solubility was 0.95 (at 30°C).

Samples that were exposed to solutions with a relative solubility of 0.6 had hardness readings that were significantly different from the controls in the first study (68 days) but not in the second study (77 days). We were not certain why the results for the two studies differed. We thought that if the unsoftened samples had been given

more time that softening might have occurred, or it might have been due to the heterogeneous nature of PVC. We noted that at the lower relative solubilities, the test pieces continued to gain weight, and therefore equilibrium was not reached. We concluded that a longer-term study should resolve the lowest relative solubility that can soften PVC.

Studies on aqueous solutions that contain several organic solutes

One question that has not been addressed in the literature is what effect does an aqueous solution that contains several different PVC solvents or swelling agents have on PVC. This is an important question for the groundwater monitoring industry since most contaminated groundwater contains several organic contaminants, each at ppb or ppm levels. In cases where there is only one organic solute in a solution, there can be interactions between the organic and the polymer, the organic and water, and among the organic molecules. In cases where there is more than one organic solute, there can be all those interactions as well as interactions among the various organic solutes (i.e., a and b, b and c, a and c, etc.). We have conducted several studies that have looked at softening by solutions containing multiple organic solutes.

In the first study, we (Parker and Ranney 1994) found that after two days' contact, the hardness readings of samples exposed to an aqueous solution that contained three PVC solvents or swelling agents were significantly lower than those of the controls, and these samples were rubbery. The test solution contained methylene chloride, 1,1,2-trichloroethane, and 1,2-dichloroethane, each with a relative solubility of 0.3. The χ values for these three compounds are reported to be less than 0.6 (Table 1). At this relative solubility, Berens would predict that none of these organic solutes should be able to soften PVC by themselves. Our own studies (Parker and Ranney 1994) have shown that methylene chloride was able to cause softening at a similar relative solubility (0.4), but it took eight weeks. The test solution in this study behaved more like a test solution containing methylene chloride at a relative solubility of 0.8, which softened the PVC pieces almost immediately (Parker and Ranney 1994).

In a follow-up study (Parker and Ranney 1995), an aqueous solution that contained 18 PVC solvents and swelling agents, each with a relative solubility of 0.05, softened PVC. There was a sig-

nificant change in the Barcol readings after only one day, and after 14 days several of the test pieces were noticeably softened when tested manually, although none actually became pliable. It is clear from these studies that there is a cumulative or possibly synergistic effect when several PVC solvents and/or swelling agents are present in an aqueous solution.

We thought that the sum of the relative solubilities (or activities) of the organic solutes in the test solution might prove to be useful in estimating the ability of a solution to soften PVC. In this study (Parker and Ranney 1995), the sum of the relative solubilities of the organic solutes was 0.9. This solution was more aggressive than a solution containing a PVC swelling agent (the TCE solution) with a relative solubility of 1.0, but less aggressive than a solution containing a PVC solvent (the methylene chloride solution) with a relative solubility of 0.8 (Parker and Ranney 1994). Although it does not explain how the organic solutes interact with the polymer, water, or each other, the sum of the relative solubilities does give some measure of the ability of an aqueous solution containing two or more organic solutes to degrade PVC.

We concluded that there were a number of unanswered questions, such as 1) whether a mixture of organic solutes that are at even lower relative solubilities could soften PVC, 2) whether PVC solvents and swelling agents have any synergistic effect on other, poorer swelling agents, and 3) what happens in solutions that contain a number of solutes with varying but very low relative solubilities?

Aqueous solutions containing organics that are miscible in water

Another issue that had not been addressed in the literature prior to our studies was whether aqueous solutions of PVC solvents and swelling agents, which are also completely miscible in water, can soften PVC. It is impossible to calculate the relative solubility of these solutions, because this involves dividing the analyte's concentration by an infinitely large number. We have conducted several studies that look at this issue.

In our most recent paper (Parker and Ranney 1995), PVC test pieces were exposed to varying concentrations of aqueous solutions containing one of four organic chemicals that were totally miscible in water: acetone, tetrahydrofuran, dimethylformamide, and pyridine. Acetone is a good swelling agent of PVC ($\chi = 0.61$ [Berens

1985]). There are no χ values available for the other three chemicals, but tests by our laboratory indicate that they are all PVC solvents (App. A, Ranney and Parker 1995). In fact, tetrahydrofuran is commonly used as a component of PVC glues.

We (Parker and Ranney 1995) found that the test solutions containing tetrahydrofuran were the most aggressive. After seven days, a 60% solution (8.4 M) dissolved the PVC pieces. By the end of a 20-week study that used lower concentrations, the hardness readings of the samples exposed to the lowest concentration tested (0.01% or 0.0014 M) were significantly lower than the controls. However, none of these samples were noticeably softened when tested manually, even the samples exposed to the highest concentration tested (10% or 1.4 M). Although there is a lot of variability in the weight change data, by the end of the study there was no additional weight gain in the samples that were exposed to the 1.0, 0.1, and 0.01% solutions. This may indicate that these solutions have reached equilibrium. If this is the case, then we would not expect any additional softening to occur.

The acetone solution was the next most aggressive solution (Parker and Ranney 1995). After seven days, the hardness readings for samples exposed to the 60% (10.2 M) solutions were significantly lower than the controls, and the samples were pliable when tested manually (i.e., could be bent in half easily). In a longer-term study (20 weeks) (Parker and Ranney 1995), there was no significant change in the hardness readings of samples that were exposed to the 2.5% (0.43 M) and 5% (0.86 M) solutions. There was a significant change in the Barcol readings of samples exposed to the 10% (1.7 M), 20% (3.4 M), and 40% (6.9 M) solutions. However, none of these samples were noticeably softened when tested manually. Although the weight gain data were highly variable, it did appear that the samples exposed to the 40%, 10%, and the 5% solutions had reached equilibrium. If this is the case, we would not expect further softening to occur. However, this is unproven at this time.

Our initial study (Parker and Ranney 1995) showed that aqueous solutions of pyridine were not highly aggressive. Even the 60% solution had no effect on the hardness readings after seven days. However, after a 20-week study (Parker and Ranney 1995), there was a significant change in the hardness readings for the samples exposed to the lowest concentration tested (20% or 2.5 M). Samples exposed to the highest concentration

tested (80% or 10 M) were actually pliable, i.e., easily bent in half.

The dimethylformamide solution was the least aggressive solution. During a 20-week study (Parker and Ranney 1995), only the test pieces exposed to the 60 (8.2 M) and 80 percent (10.9 M) solutions had hardness readings that were significantly lower than the controls. The samples that were exposed to the 80% solution were also pliable. Because there was no additional weight gain in any of these samples during the last four weeks of the study, equilibrium may have been reached.

It is clear from this study that PVC solvents and swelling agents that are completely miscible in water can soften PVC when mixed with water. These results are puzzling in that the aqueous solutions of acetone were more aggressive than the aqueous solutions of the dimethylformamide and pyridine, and yet our laboratory studies show that these two chemicals are more aggressive than acetone (App. A, Ranney and Parker 1995). Therefore, we would expect that the aqueous acetone solutions would be the least aggressive solutions. It appears that the ability of aqueous solutions of these chemicals does not completely correlate with the neat chemicals' ability to dissolve or swell PVC. This can be related to other interactions that occur in aqueous solution. According to Leggett (pers. comm. 1995*) acetone is less basic than the other solutes and would interact less strongly with water, and thus an aqueous solution would have a higher activity on a molar basis.

PURPOSE OF THIS STUDY

The purpose of this study was to determine:

- 1) whether softening progresses further in those solutions already showing a significant change in hardness readings but are not yet pliable?
- 2) for solutions with single organic solutes, what is the lowest relative solubility where softening occurs?
- 3) for solutions with multiple organic solutes, what is the lowest relative solubility where softening occurs?
- 4) for solutions with multiple organic solutes, how well does the sum of the relative solubilities of the solutes in the solution correlate with the solution's ability to soften PVC?

* D. Leggett, Research Chemist, USA Cold Regions Research and Engineering Laboratory, Hanover, New Hampshire.

5) what occurs in an aqueous solution that contains a number of PVC solvents and swelling agents at varying relative solubilities?

6) in solutions that contain several organic solutes, do solutes that are solvents and swelling agents of PVC have some type of synergistic effect on poorer swelling agents?

7) what is the lowest concentration of acetone in aqueous solution that can soften PVC?

8) what occurs when PVC solvents and swelling agents that are miscible in water are also present in an aqueous solution with other PVC solvents and swelling agents?

MATERIALS AND METHODS

Small pieces of PVC were cut from 2-in.- (5.1 cm) diameter, schedule 40 well casing. The overall dimensions of the pieces were approximately $9 \times 9 \times 1.1$ mm (L \times W \times H), giving a total surface area of approximately 2.1 cm² and approximate weight of 0.12 g. Special care was taken to eliminate contamination from grease or oil during the cutting procedure. The pieces of PVC were washed by placing them in a glass beaker containing a 2% solution of Liquinox detergent and deionized (Millipore) water for 30 minutes, and stirring them intermittently with a glass rod for approximately five minutes. The pieces were then rinsed with copious amounts of deionized water to ensure complete removal of the detergent, and were then spread on paper towels, blotted, and left to air dry.

All the test solutions were made up by adding appropriate volumes of the neat analyte (based on the density of the compound) to deionized water in volumetric flasks. To prevent any losses due to biological activity, mercuric chloride was added to the test solution so that the final concentration was 40 mg/L. A glass-covered magnetic stirring bar was added to each flask, and the flasks were then filled with deionized water to within ~1 mL of the top. This left a minimal head space once the glass stopper was inserted. The stoppers were then wrapped tightly with parafilm to reduce vapor loss and the solutions were mixed for a minimum of three days on magnetic stirring plates.

After mixing for several days, the solutions were checked with a magnifying glass to ensure that there were no undissolved droplets of solvent visible. Once it was determined that the organic(s) had dissolved, pre-weighed PVC pieces

were added to 22-mL borosilicate glass vials. The vials were then filled to capacity with the test solution so there was no headspace, and capped with Teflon-lined plastic caps. PVC pieces added to deionized water served as controls. In all the studies, there were five replicates for each type of solution (relative solubility) and time. The test samples were stored in the dark at room temperature. The ratio of surface area of PVC to solution volume was 0.079 cm²/mL, and the ratio of solution volume to volume of PVC was approximately 310:1.

On the sampling day, the pieces of PVC were removed from the solution, placed on paper towels, blotted, and allowed to air dry for one minute before weighing (± 0.0001 g). In all cases, the pieces of PVC were handled using stainless steel forceps. Immediately after weighing, a Model 935 Barcol Impressor was used to measure the hardness of the PVC. A description on the use of the Barcol Impressor on rigid plastics can be found in ASTM Standard Test Method D 2583-87 (ASTM 1988). However, the impressor employed in this study is made for indenting softer plastics and is not the model described in the ASTM test method. This model was selected so that measurements could be taken during the softening process. The flexibility was also tested by holding the pieces with forceps and noting the amount of force necessary to bend the sample.

In all these studies, Analysis of Variance and Fisher's Protected Least Significant Difference (LSD) tests were performed on the Barcol readings and weights of the samples.

Long-term methylene chloride study

In this study, the PVC test pieces were exposed to aqueous solutions of methylene chloride with the following relative solubilities: 0.01, 0.05, 0.1, 0.2, and 0.4. Contact times were 3, 6, 9, 12, 15, and 18 months. The test solutions were changed every six months. Relative solubilities were based upon a solubility of 19,000 mg/L. Because Montgomery and Welkom (1990) listed a range of solubility values, we conducted our own solubility tests and then selected the published value that most closely agreed with our findings (Parker and Ranney 1994).

Long-term TCE study

In this study, PVC was exposed to aqueous solutions of TCE at five relative solubilities: 0.05, 0.10, 0.20, 0.40, and 0.60. Contact times were the

Table 2. Relative solubility and molarity of PVC solvent and swelling agents used in the test solutions.

Solvent	Solubility (mg/L)	Mol. wt. (g/mole)	Solution A		Solution B	
			Rel. solubility	Molarity	Rel. solubility	Molarity
bromodichloromethane	4,500 ¹	163.8 ¹	0.08	0.022	0.07	0.019
chlorobenzene	500 ¹	112.56 ¹	0.04	0.002	0.10	0.004
chloroform	8,000 ¹	119.38 ¹	0.12	0.080	0.09	0.060
cyclohexanone	23,000 ²	98.2 ²	0.09	0.211	0.12	0.281
1,2-dichloroethane	8,690 ¹	98.96 ¹	0.15	0.132	0.08	0.070
1,1,2-trichloroethane	4,500 ¹	133.41 ¹	0.10	0.034	0.04	0.013
methylene chloride	19,000 ¹	84.93 ¹	0.05	0.112	0.02	0.045
methyl isobutyl ketone	17,000 ¹	100.2 ¹	0.07	0.119	0.08	0.136
nitrobenzene	1,900 ¹	123.1 ¹	0.08	0.012	0.15	0.023
<i>m</i> -nitrotoluene	500 ¹	137.13 ¹	0.02	0.001	0.05	0.002
Total				0.725		0.653

¹ Value taken from Montgomery and Welkom (1990).

² Value taken from Montgomery (1991).

same as the previous study, and the test solutions were changed every six months. An aqueous solubility of 1000 mg/L (Dean 1985) was used.

Studies on aqueous solutions that contain several organic solutes

Short-term study

In this study, we exposed PVC to two solutions containing ten PVC solvents and swelling agents (Table 2). Table 2 lists the aqueous solubility values we used to determine relative solubility. The sum of the relative solubilities of the organic solutes in the solutions was 0.80, the relative solubility of each analyte varied from 0.02 to 0.15, and the relative solubility of each analyte varied from solution A to solution B. Contact times were 8, 24, 72, and 168 hours (seven days).

Mid-term study

The test pieces of PVC were exposed to a solution of either TCE, methylene chloride, or a combination of TCE and methylene chloride. The relative solubilities for methylene chloride were either 0 or 0.4. The relative solubilities for the TCE were either 0, 0.2, 0.4, or 0.6. Contact times were 2, 7, 14, 28, and 56 days.

First long-term study

PVC test pieces were exposed to two aqueous solutions of 18 organic chemicals (Table 3) that we have found to be either solvents or swelling agents of PVC (App. A, Ranney and Parker 1995).

The solubilities of these 18 chemicals are also given in Table 3. The protocol for this experiment differed from our other experiments in that smaller test pieces ($\sim 9 \times 9 \times 0.9$ mm) that were left over from our previous studies (Parker and Ranney 1994, 1995) were used. These dimensions made the total surface area 1.9 cm²,

Table 3. Listing of 18 PVC solvent and swelling agents used in studies and their aqueous solubilities.

	Solubility (mg/L)
benzotrile	2,000 ³
bromochloromethane	16,691 ²
bromodichloromethane	4,500 ¹
bromoform	3,010 ¹
chlorobenzene	500 ¹
chloroform	8,000 ¹
cyclohexanone	23,000 ²
dibromochloromethane	4,000 ¹
1,2-dibromoethane	4,300 ³
1,2-dichloroethane	8,690 ¹
<i>cis</i> -1,2-dichloroethylene	7,700 ³
<i>trans</i> -1,2-dichloroethylene	6,300 ³
methylene chloride	19,000 ¹
methyl isobutyl ketone	17,000 ¹
nitrobenzene	1,900 ¹
<i>m</i> -nitrotoluene	500 ¹
<i>o</i> -nitrotoluene	600 ²
1,1,2-trichloroethane	4,500 ¹

¹ Value taken from Montgomery and Welkom (1990).

² Value taken from Montgomery (1991).

³ Value taken from Dean (1985).

the approximate weight 0.09 g, the ratio of surface area to solution volume 0.076, and the solution volume to PVC volume approximately 386:1. Another difference with this experiment was that the Teflon-lined caps were lined with aluminum foil to prevent any losses due to sorption by caps, and as a result mercuric chloride was not added to the test solutions because it would degrade the aluminum foil. The relative solubility of each of the analytes was either 0.010 or 0.020. The sum of the relative solubilities of the organic solutes was either 0.18 or 0.36, respectively. Samples were taken after 20 and 40 weeks' contact.

Second long-term study

In this study, pieces of PVC were exposed to four aqueous solutions of the same eighteen PVC solvents and swelling agents (Table 3). This time the relative solubility of the analytes was either 0.001, 0.005, 0.010, or 0.020. This makes the sum of the relative solubilities of the organic solutes either 0.018, 0.09, 0.18, or 0.36, respectively. Samples were taken after 3, 6, 9, 12, 15, and 18 months, and the test solutions were changed every six months.

Studies on aqueous solutions containing organics that are miscible in water

Long-term acetone study

In this study, pieces of PVC were exposed to aqueous solutions that contained the following percent acetone: 0.5%, 1%, 5%, 10%, and 20% (w/v). Contact times were again 3, 6, 9, 12, 15, and 18 months. The test solutions were changed every six months.

Short-term study with acetone and methylene chloride

PVC was exposed to different solutions containing methylene chloride and acetone. The relative solubility of the methylene chloride was 0.4, while the concentration of acetone in the solution was either 0%, 0.5%, 1.0%, 5.0%, 10%, or 20% (w/v). Contact times were 2, 4, 7, 14, and 21 days.

RESULTS AND DISCUSSION

Long-term methylene chloride studies

Table 4 gives the hardness readings and percent weight gain data for samples exposed to aqueous solutions with relative solubilities of 0.01, 0.05, 0.10, 0.20, and 0.40. After 18 months' exposure, there was no significant change in the hardness readings of samples that were exposed to aqueous solutions of methylene chloride with relative solubilities that were less than 0.1. With one ex-

Table 4a. Hardness (Barcol) readings of PVC exposed to methylene chloride solutions.

Relative solubility	Contact time, months					
	3	6	9	12	15	18
0.00	71	74	78	77	78	78
0.00	71	73	79	77	78	80
0.00	72	73	78	78	76	79
0.00	73	74	76	76	78	80
0.00	73	74	80	78	78	78
\bar{X}	72.0	73.6	78.2	77.2	77.6	79.0
0.01	73	74	78	77	78	79
0.01	72	74	79	79	77	79
0.01	71	72	80	76	78	80
0.01	70	73	77	78	79	79
0.01	74	73	78	77	77	78
\bar{X}	72.0	73.2	78.4	77.4	77.8	79.0
0.05	72	74	76	77	76	77
0.05	72	74	78	78	79	80
0.05	71	73	79	76	78	79
0.05	70	73	76	76	76	78
0.05	70	72	77	78	77	78
\bar{X}	71.0	73.2	77.2	77.0	77.2	78.4
0.10	71	74	76	76	76	77
0.10	72	74	79	74	78	76
0.10	71	74	77	77	75	78
0.10	72	73	78	76	76	77
0.10	71	73	80	75	74	75
\bar{X}	71.4	73.6	78.0	75.6*	75.8*	76.6*
0.20	71	73	77	75	73	77
0.20	71	74	76	76	74	76
0.20	70	72	77	75	73	74
0.20	71	73	76	73	74	74
0.20	72	73	74	77	73	73
\bar{X}	71.0	73.0	76.0*	75.2*	73.4*	74.8*
0.40	69	72	74	74	73	71
0.40	70	72	77	74	72	72
0.40	71	71	76	75	71	74
0.40	70	72	76	75	71	74
0.40	69	73	76	73	72	73
\bar{X}	69.8*	72.0*	75.8*	74.2*	71.8*	72.8*

*Significantly different from controls.

ception, there was also no significant change in the weight change data. These results support our hypothesis that softening will not be detected at activities where diffusion is reported to be slowest (Fickian), i.e., for PVC solvents and better swelling agents less than 0.1 (Vonk 1985, 1986).

Table 4b. Percent weight gain of PVC exposed to methylene chloride solutions.

Relative solubility	Contact time, months					
	3	6	9	12	15	18
0.00	0.13	0.24	0.26	0.20	0.33	0.17
0.00	0.26	0.28	0.17	0.09	0.00	0.32
0.00	0.15	0.35	0.35	0.28	0.25	0.19
0.00	0.17	0.35	0.26	0.16	0.27	0.16
0.00	0.25	0.35	0.27	0.14	0.20	0.26
\bar{X}	0.19 ^a	0.31 ^{a,b}	0.26 ^a	0.17 ^a	0.21 ^a	0.22 ^a
S.D.	0.06	0.05	0.07	0.07	0.13	0.07
0.01	0.35	0.18	0.26	0.26	0.35	0.17
0.01	0.17	0.26	0.35	0.34	0.17	0.24
0.01	0.25	0.27	0.41	0.36	0.16	0.21
0.01	0.34	0.36	0.35	0.27	0.08	0.18
0.01	0.41	0.42	0.42	0.35	0.34	0.08
\bar{X}	0.30 ^a	0.30 ^a	0.36 ^a	0.32 ^{a,b}	0.22 ^a	0.18 ^a
S.D.	0.09	0.09	0.06	0.05	0.12	0.06
0.05	0.44	0.25	0.40	0.34	0.27	0.35
0.05	0.34	0.33	0.40	0.34	0.33	0.50
0.05	0.34	0.44	0.36	0.31	0.24	0.17
0.05	0.34	0.44	0.33	0.61	0.33	0.41
0.05	0.24	0.25	0.34	0.34	0.25	0.28
\bar{X}	0.34 ^a	0.34 ^{a,b}	0.37 ^a	0.39 ^{b,c}	0.28 ^a	0.34 ^{a,b}
S.D.	0.07	0.09	0.03	0.12	0.04	0.13
0.10	0.33	0.41	0.62	0.53	0.72	0.43
0.10	0.37	0.41	0.55	0.42	0.52	0.64
0.10	0.43	0.62	0.53	0.50	0.59	0.52
0.10	0.26	0.69	0.64	0.54	0.60	0.58
0.10	0.32	0.37	0.69	0.69	0.51	0.44
\bar{X}	0.34 ^a	0.50 ^b	0.60 ^b	0.54 ^{c,d}	0.59 ^b	0.52 ^{b,c}
S.D.	0.06	0.14	0.06	0.10	0.08	0.09
0.20	0.75	1.01	0.74	0.74	0.82	0.60
0.20	0.72	0.64	1.11	0.59	0.87	0.67
0.20	0.87	0.76	0.66	0.61	0.92	0.58
0.20	0.61	0.60	1.00	0.81	1.05	0.61
0.20	0.65	0.75	0.83	0.85	0.84	0.59
\bar{X}	0.72 ^b	0.75 ^c	0.87 ^c	0.72 ^d	0.90 ^c	0.61 ^c
S.D.	0.10	0.16	0.19	0.12	0.09	0.03
0.40	1.33	1.39	1.69	2.38	2.37	2.18
0.40	1.51	0.79	1.91	2.15	2.13	2.01
0.40	1.65	1.01	1.66	1.94	1.67	2.20
0.40	1.99	1.25	1.36	2.26	2.50	1.41
0.40	1.43	1.39	1.66	1.67	2.01	1.97
\bar{X}	1.58 ^c	1.17 ^d	1.66 ^d	2.08 ^e	2.14 ^d	1.95 ^d
S.D.	0.26	0.26	0.19	0.28	0.32	0.32

S.D. = standard deviation

Values with different letters, taken at the same time, are significantly different.

Hardness readings of samples that were exposed to aqueous solutions with a relative solubility of 0.1 were significantly lower than the controls after one year. Samples exposed to solutions with higher relative solubilities (0.2, 0.4) softened earlier in the study. However, none of these samples were pliable. Presumably these materials become slightly softened but never become pliable or rubbery. This conclusion is supported by the fact that there does not appear to be much additional weight gain after the first three to six months for any of the samples exposed to the solutions with relative solubilities of 0.10 and 0.20 and after 12 months in samples exposed to solutions with a relative solubility of 0.40. This may mean that the test pieces have reached equilibrium with the test solution, and thus we do not expect any further changes.

We also observed that softening occurred more slowly for the samples exposed to solutions with relative solubilities of 0.10 and 0.20 than what we observed previously (Parker and Ranney 1994). This is most likely a result of the pieces being approximately 20% thicker in this study.

Long-term TCE studies

Table 5 gives the results of exposure to aqueous solutions of TCE with relative solubilities of 0.05, 0.1, 0.2, 0.4, and 0.6. The samples exposed to the solutions with relative solubilities of 0.6 and 0.4 had significant changes in their hardness readings; however, they did not become pliable. It should be noted that TCE has never before been shown to be active at such a low relative solubility, and that it took more than a year before significant changes in hardness were detected in those samples. There were also significant changes in the weights of the samples exposed to the solutions with relative solubilities of 0.4 and 0.6.

It is clear from this and the previous study that small but measurable changes occur in the weight

Table 5a. Hardness (Barcol) readings of PCE exposed to TCE solutions.

Relative solubility	Contact time, months					
	3	6	9	12	15	18
0.00	70	74	77	76	80	80
0.00	72	73	78	77	78	77
0.00	71	74	80	78	79	78
0.00	71	72	77	74	78	79
0.00	71	71	77	75	78	79
\bar{X}	71.0	72.8	77.8	76.0	78.6	78.6
0.05	70	73	74	76	78	79
0.05	71	74	75	75	79	78
0.05	71	72	77	75	78	79
0.05	71	72	78	78	79	78
0.05	72	73	79	74	78	78
\bar{X}	71.0	72.8	76.6	75.6	78.4	78.4
0.10	70	74	77	76	79	79
0.10	71	74	76	74	76	80
0.10	70	72	73	77	77	78
0.10	71	72	80	78	78	78
0.10	69	73	80	77	78	78
\bar{X}	70.2	73.0	77.2	76.4	77.6	78.6
0.20	71	72	75	78	77	77
0.20	71	73	76	74	78	77
0.20	69	73	74	75	79	79
0.20	71	74	79	75	76	79
0.20	72	72	75	74	77	78
\bar{X}	70.8	72.8	75.8	75.2	77.4	78.0
0.40	71	72	77	74	77	76
0.40	71	73	75	76	76	78
0.40	71	74	74	75	76	76
0.40	69	72	78	76	74	76
0.40	70	73	73	74	74	78
\bar{X}	70.4	72.8	75.4	75.0	75.4*	76.8*
0.60	68	71	75	73	75	76
0.60	70	70	73	72	75	75
0.60	70	71	74	73	74	77
0.60	68	70	75	74	75	75
0.60	69	72	72	75	73	77
\bar{X}	69.0*	70.8*	73.8*	73.4*	74.4*	76.0*

*Significantly different from controls.

and hardness of samples exposed to aqueous solutions of PVC solvents and swelling agents, and that these changes occur at lower relative solubilities than was predicted by Berens.

Studies on aqueous solutions that contain several organic solutes

Short-term study

In this study, we exposed PVC to two solutions containing ten organic chemicals that were

either a PVC solvent or swelling agent (Table 2). The ten organic chemicals were selected based on published χ values (Berens 1985, Vonk 1985) and on our own tests, where small pieces of PVC casing were subjected to neat organic chemicals (App. A, Ranney and Parker 1995). The sum of the relative solubilities of the organic chemicals in the two solutions was the same, 0.80. However, the relative solubility of each analyte varied from 0.02 to 0.15 and the relative solubility of each analyte varied from solution A to solution B.

Both solutions caused rapid (within one day)

Table 5b. Percent weight gain of PVC exposed to TCE solutions.

Relative solubility	Contact time, months					
	3	6	9	12	15	18
0.00	0.26	0.00	0.26	0.17	0.27	0.17
0.00	0.23	0.25	0.25	0.21	0.17	0.25
0.00	0.25	0.25	0.33	0.33	0.34	0.09
0.00	0.17	0.39	0.08	0.25	0.16	0.18
0.00	0.26	0.35	0.33	0.16	0.08	0.33
\bar{X}	0.23 ^a	0.25 ^a	0.25 ^a	0.22 ^a	0.20 ^a	0.20 ^a
S.D.	0.04	0.15	0.10	0.07	0.10	0.09
0.05	0.25	0.25	0.25	0.17	0.36	0.16
0.05	0.32	0.32	0.26	0.34	0.26	0.19
0.05	0.08	0.42	0.17	0.26	0.27	0.16
0.05	0.18	0.33	0.17	0.33	0.26	0.34
0.05	0.20	0.26	0.27	0.25	0.42	0.24
\bar{X}	0.21 ^a	0.32 ^a	0.22 ^a	0.27 ^{a,b}	0.31 ^b	0.22 ^a
S.D.	0.09	0.07	0.05	0.07	0.07	0.07
0.10	0.26	0.29	0.25	0.25	0.33	0.09
0.10	0.25	0.16	0.17	0.18	0.26	0.20
0.10	0.18	0.29	0.34	0.27	0.35	0.16
0.10	0.25	0.39	0.51	0.40	0.17	0.34
0.10	0.18	0.33	0.18	0.17	0.17	0.34
\bar{X}	0.22 ^a	0.29 ^a	0.29 ^{a,b}	0.25 ^{a,b}	0.26 ^{a,b}	0.23 ^a
S.D.	0.04	0.08	0.14	0.09	0.08	0.11
0.20	0.34	0.41	0.00	0.35	0.34	0.28
0.20	0.35	0.24	0.27	0.24	0.42	0.25
0.20	0.35	0.34	0.34	0.41	0.25	0.27
0.20	0.26	0.25	0.36	0.34	0.35	0.34
0.20	0.34	0.25	0.17	0.41	0.18	0.34
\bar{X}	0.33 ^b	0.30 ^a	0.23 ^a	0.35 ^b	0.31 ^{a,b}	0.30 ^{a,b}
S.D.	0.04	0.08	0.15	0.07	0.09	0.04
0.40	0.33	0.30	0.43	0.40	0.53	0.57
0.40	0.33	0.41	0.42	0.44	0.44	0.33
0.40	0.41	0.41	0.41	0.52	0.40	0.34
0.40	0.40	0.34	0.43	0.47	0.33	0.25
0.40	0.32	0.27	0.29	0.59	0.54	0.34
\bar{X}	0.36 ^b	0.35 ^a	0.40 ^b	0.49 ^c	0.45 ^c	0.37 ^b
S.D.	0.04	0.07	0.06	0.07	0.09	0.12
0.60	0.75	0.65	0.66	0.94	0.86	0.83
0.60	0.57	0.72	0.88	0.90	0.75	0.72
0.60	0.61	0.49	0.65	0.90	0.78	0.55
0.60	0.60	0.43	0.74	0.78	0.82	0.63
0.60	0.66	0.68	0.75	0.78	0.74	0.53
\bar{X}	0.64 ^c	0.60 ^b	0.74 ^c	0.86 ^d	0.79 ^d	0.65 ^c
S.D.	0.07	0.12	0.09	0.07	0.05	0.13

S.D. = standard deviation

Values with different letters, taken at the same time, are significantly different.

changes in the hardness readings of PVC, although none of these samples actually became pliable (Table 6). The ability of the solution to soften PVC appears to be correlated with the sum of the relative solubilities of the organic solutes, since the behavior of the two test solutions was between that of a methylene chloride solution (i.e., a solution with a PVC solvent) with a relative solubility

of 0.8 and a TCE solution (a solution with a swelling agent) with a relative solubility of 0.8. There does not appear to be a synergistic effect with this solution, i.e., the ability to soften PVC does not appear to be enhanced beyond a simple cumulative effect.

However, the data (both weight change and hardness values) also show the limitations of us-

Table 6a. Hardness (Barcol) readings of PVC exposed to aqueous solutions of several organic solutes (short-term study).

Treatment	Contact time, hours			
	8	24	72	168
control	73	77	74	74
control	69	74	71	72
control	67	76	70	69
control	71	77	74	72
control	68	75	70	71
\bar{X}	69.6 ^a	75.8 ^a	71.8 ^a	71.6 ^a
solution A	63	69	58	60
solution A	64	67	62	64
solution A	65	72	62	62
solution A	69	70	65	64
solution A	64	67	64	60
\bar{X}	65.0 ^b	69.0 ^b	62.2 ^b	62.0 ^c
solution B	67	70	63	68
solution B	69	71	68	64
solution B	71	71	64	66
solution B	65	66	69	66
solution B	69	68	64	65
\bar{X}	68.2 ^{a,b}	69.2 ^b	65.6 ^b	65.8 ^b

Values with different letters, taken at the same time, are significantly different.

ing this approach (i.e., summing the relative solubilities), since solution A appears to be more aggressive than solution B (Table 6), even though they both have the same cumulative relative solubilities. The reason solution A was more aggressive can be explained by the fact that the sum of the mole fractions of the organic solutes was greater in solution A (Table 2). In addition, the sum of the mole fractions of the four most aggressive chemicals (cyclohexanone, 1,2-dichloroethane, methylene chloride, and nitrobenzene) (App. A, Ranney and Parker 1995) in the test solution was also greater.

Mid-term study

Test pieces of PVC that were exposed to a TCE solution with a relative solubility of 0.2 were not softened during the course of this study (56 days) (Table 7). However, samples exposed to a solution containing TCE with a relative solubility of 0.2 and methylene chloride at a relative solubility of 0.4 were softened after seven days. Again, none of these samples were pliable. This test solution was less aggressive than we (Parker and Ranney 1994) found a methylene chloride solution with a relative solubility of 0.6 (i.e., a solution contain-

Table 6b. Percent weight gain of PVC exposed to aqueous solutions of several organic solutes (short-term study).

Treatment	Contact time, hours			
	8	24	72	168
control	0.00	0.00	0.23	0.24
control	0.08	0.08	0.08	0.40
control	0.15	0.25	0.08	0.43
control	0.09	0.16	0.25	0.27
control	0.08	0.08	0.17	0.42
\bar{X}	0.08 ^a	0.11 ^a	0.16 ^a	0.35 ^a
S.D.	0.05	0.09	0.08	0.09
solution A	1.23	2.35	3.36	5.58
solution A	1.07	2.57	3.58	5.29
solution A	1.03	2.48	3.90	5.29
solution A	1.08	2.34	3.61	5.00
solution A	1.07	2.36	3.70	5.67
\bar{X}	1.10 ^c	2.42 ^c	3.63 ^c	5.37 ^c
S.D.	0.08	0.10	0.19	0.27
solution B	0.65	1.65	2.55	3.44
solution B	0.79	1.59	2.27	3.54
solution B	0.82	1.58	2.62	3.74
solution B	0.73	1.73	2.36	3.53
solution B	0.78	1.42	2.60	4.20
\bar{X}	0.75 ^b	1.59 ^b	2.48 ^b	3.69 ^b
S.D.	0.07	0.11	0.16	0.31

S.D. = standard deviation

Values with different letters, taken at the same time, are significantly different.

ing a PVC solvent) to be, but more aggressive than we (Parker and Ranney 1994) found a TCE solution with the same relative solubility (0.6) to be. Thus, there does not appear to be an enhanced or synergistic effect of the organic solutes in this test solution, and taking the sum of the relative solubilities was useful in roughly predicting the ability of this test solution to soften PVC.

First long-term study

PVC pieces were exposed to two aqueous solutions that contained 18 organic chemicals that were either a PVC solvent or swelling agent (Table 8). In these solutions, the relative solubility of each chemical was either 0.01 or 0.02, and the sums of the relative solubilities of all the organic solutes in the test solutions were 0.18 and 0.36, respectively. After 40 weeks' exposure, the test pieces exposed to the test solution with the higher total relative solubility (0.36) had significant changes in hardness readings (but were not pliable), while those exposed to the solution with the lower total relative solubility (0.18) did not. However, there was a significant increase in the weight of the test pieces exposed to this test solution when compared with the controls.

Table 7a. Hardness (Barcol) readings of PVC exposed to aqueous solutions of methylene chloride (MC) and TCE.

Relative solubility MC, TCE	Contact time, days				
	2	7	14	28	56
control	75	70	73	79	76
control	73	71	74	76	74
control	74	73	75	78	75
control	73	74	72	77	75
control	76	75	73	76	77
\bar{X}	74.2	72.6	73.4	77.2	75.4
0, 0.2	72	73	74	78	76
0, 0.2	74	75	73	77	75
0, 0.2	75	74	71	72	74
0, 0.2	74	72	70	76	75
0, 0.2	73	71	72	78	74
\bar{X}	73.6	73.0	72.0	76.2	74.8
0, 0.4	74	69	74	74	76
0, 0.4	73	70	74	78	75
0, 0.4	74	73	74	78	75
0, 0.4	74	71	73	73	73
0, 0.4	73	72	70	78	75
\bar{X}	73.6	71.0	73.0	76.2	74.8
0, 0.6	73	73	72	77	74
0, 0.6	75	71	73	75	75
0, 0.6	76	72	68	79	74
0, 0.6	74	68	70	78	73
0, 0.6	75	71	73	77	74
\bar{X}	74.6	71.0	71.2	77.2	74.0 *
0.4, 0	76	68	75	74	74
0.4, 0	72	72	70	78	73
0.4, 0	74	73	71	76	72
0.4, 0	73	71	69	78	73
0.4, 0	75	71	72	73	72
\bar{X}	74.0	71.0	71.4	75.8	72.8*
0.4, 0.2	73	66	70	74	72
0.4, 0.2	76	67	69	73	70
0.4, 0.2	74	66	72	75	72
0.4, 0.2	75	66	70	72	71
0.4, 0.2	72	69	72	74	72
\bar{X}	74.0	66.8*	70.6*	73.6*	71.4*
0.4, 0.4	73	70	70	73	70
0.4, 0.4	71	71	68	71	70
0.4, 0.4	74	68	68	74	68
0.4, 0.4	73	65	71	75	70
0.4, 0.4	74	66	69	70	69
\bar{X}	73.0	68.0*	69.2*	72.6*	69.4*
0.4, 0.6	66	63	60	67	62
0.4, 0.6	70	64	63	63	62
0.4, 0.6	70	65	65	70	65
0.4, 0.6	68	65	64	64	61
0.4, 0.6	71	67	61	62	63
\bar{X}	69.0*	64.8*	62.6*	65.2*	62.6*

*Significantly different from controls.

Table 7b. Percent weight gain of PVC exposed to aqueous solutions of methylene chloride (MC) and TCE.

Relative solubility MC, TCE	Contact time, days				
	2	7	14	28	56
control	0.08	0.16	0.16	0.00	0.25
control	0.08	0.08	0.08	0.16	0.17
control	0.40	0.17	0.08	0.17	0.09
control	0.00	0.17	0.09	0.25	0.16
control	0.09	0.16	0.00	0.16	0.16
\bar{X}	0.13 ^a	0.15 ^a	0.08 ^a	0.15 ^a	0.17 ^a
S.D.	0.15	0.04	0.06	0.09	0.06
0, 0.2	0.08	0.25	0.24	0.24	0.27
0, 0.2	0.08	0.16	0.25	0.24	0.23
0, 0.2	0.00	0.23	0.08	0.08	0.25
0, 0.2	0.00	0.26	0.24	0.24	0.24
0, 0.2	0.00	0.16	0.16	0.16	0.18
\bar{X}	0.03 ^a	0.21 ^a	0.19 ^a	0.19 ^a	0.23 ^a
S.D.	0.04	0.05	0.08	0.07	0.03
0, 0.4	0.16	0.24	0.24	0.24	0.32
0, 0.4	0.08	0.24	0.26	0.31	0.44
0, 0.4	0.18	0.17	0.24	0.17	0.41
0, 0.4	0.08	0.26	0.23	0.34	0.28
0, 0.4	0.08	0.24	0.24	0.25	0.32
\bar{X}	0.12 ^a	0.23 ^a	0.24 ^a	0.26 ^a	0.35 ^a
S.D.	0.05	0.04	0.01	0.07	0.07
0, 0.6	0.16	0.40	0.44	0.41	0.32
0, 0.6	0.17	0.47	0.46	0.49	0.50
0, 0.6	0.16	0.34	0.49	0.46	0.32
0, 0.6	0.17	0.25	0.37	0.48	0.32
0, 0.6	0.00	0.33	0.48	0.42	0.49
\bar{X}	0.13 ^a	0.36 ^{a,b}	0.45 ^a	0.45 ^a	0.39 ^a
S.D.	0.07	0.08	0.05	0.04	0.10
0.4, 0	0.37	0.59	0.81	1.12	1.26
0.4, 0	0.33	0.74	0.84	0.93	1.42
0.4, 0	0.24	0.73	0.99	1.06	1.18
0.4, 0	0.43	0.56	0.91	1.15	1.23
0.4, 0	0.33	0.55	0.89	1.29	1.00
\bar{X}	0.34 ^b	0.63 ^b	0.89 ^b	1.11 ^b	1.22 ^b
S.D.	0.07	0.09	0.07	0.13	0.15
0.4, 0.2	0.56	0.83	1.31	1.03	1.55
0.4, 0.2	0.56	1.01	1.09	1.42	1.90
0.4, 0.2	0.49	1.05	1.26	1.71	1.69
0.4, 0.2	0.51	1.02	1.22	1.53	1.62
0.4, 0.2	0.55	1.12	1.23	1.89	1.72
\bar{X}	0.53 ^c	1.01 ^c	1.22 ^b	1.52 ^b	1.70 ^c
S.D.	0.03	0.11	0.08	0.32	0.13
0.4, 0.4	1.04	1.91	2.11	1.79	3.19
0.4, 0.4	0.89	1.80	2.18	2.63	2.80
0.4, 0.4	0.97	1.80	1.86	3.36	2.54
0.4, 0.4	0.83	1.85	2.09	2.98	2.22
0.4, 0.4	0.49	1.74	2.21	3.12	2.38
\bar{X}	0.84 ^d	1.82 ^d	2.09 ^c	2.78 ^d	2.63 ^d
S.D.	0.21	0.07	0.14	0.61	0.38
0.4, 0.6	2.72	5.02	6.88	4.49	5.78
0.4, 0.6	2.35	3.80	5.05	5.07	5.48
0.4, 0.6	2.50	5.06	5.24	5.24	5.21
0.4, 0.6	2.18	5.37	4.92	5.81	5.48
0.4, 0.6	2.36	4.05	5.67	5.97	5.25
\bar{X}	2.42 ^e	4.66 ^e	5.55 ^d	5.32 ^d	5.44 ^e
S.D.	0.20	0.69	0.80	0.60	0.23

S.D. = standard deviation

Values with different letters, taken at the same time, are significantly different.

Table 8. Hardness (Barcol) readings and percent weight gain of PVC exposed to aqueous solutions of several organic solutes (first long-term study).

Relative solubility	20 week		40 week	
	Barcol	% Δ weight	Barcol	% Δ weight
control	78	0.40	76	0.41
control	77	0.37	74	0.47
control	74	0.53	76	0.40
control	76	0.41	74	0.41
control	77	0.36	76	0.37
\bar{X}	76.4 ^a	0.41 ^a	75.2 ^a	0.41 ^a
S.D.		0.07		0.04
0.01	74	0.59	75	0.54
0.01	74	0.50	74	0.64
0.01	76	0.61	75	0.70
0.01	76	0.56	74	0.36
0.01	77	0.47	73	0.52
\bar{X}	75.4 ^a	0.55 ^b	74.2 ^{a,b}	0.55 ^b
S.D.		0.06		0.13
0.02	76	0.92	72	0.93
0.02	74	0.65	75	0.99
0.02	77	0.91	74	0.93
0.02	76	0.75	73	0.80
0.02	72	0.68	72	0.92
\bar{X}	75.0 ^a	0.78 ^c	73.2 ^b	0.91 ^c
S.D.		0.13		0.07

S.D. = standard deviation

Values with different letters, taken at the same time, are significantly different.

Second long-term study

Table 9 gives the results of exposing PVC to a test solution containing the same eighteen organic solutes at four different relative solubilities for up to 18 months. The relative solubility of each of the organic solutes was either 0.001, 0.005, 0.010, or 0.020. Thus, the sums of the relative solubilities of the organic solutes in these solutions were 0.018, 0.090, 0.18, and 0.36, respectively.

The test results show that when the sum of the relative solubilities of the organic solutes in the test solutions was less than 0.10, softening did not occur. Test pieces that were exposed to organic solutes with relative solubilities of 0.010 and 0.020 (or cumulative relative solubilities of 0.18 and 0.36) had hardness readings that were significantly lower than the controls but were not rubbery. These results agree with our previous findings for methylene chloride, i.e., softening does not occur when the relative solubility of the analyte in the test solution is less than 0.1. The weight gain data also support these conclusions. Significant changes in weight were found in samples exposed to solutions with relative solubilities of 0.01 and 0.02. Generally, there was not significant change in the weights of samples exposed to solutions with lower relative solubilities.

Table 9a. Hardness (Barcol) readings of PVC exposed to aqueous solutions of several organic solutes (second long-term study).

Relative solubility	Contact time, months					
	3	6	9	12	15	18
control	71	74	77	79	79	79
control	72	75	76	78	77	80
control	73	73	75	76	77	78
control	72	74	79	77	78	78
control	73	75	77	77	79	78
\bar{X}	72.2	74.2	76.8	77.4	78.0	78.6
0.001	71	74	77	78	77	78
0.001	73	75	76	77	79	78
0.001	71	74	75	76	80	79
0.001	72	75	75	77	78	77
0.001	73	73	76	77	78	78
\bar{X}	72.0	74.2	75.8	77.0	78.4	78.0
0.005	71	74	77	75	77	79
0.005	72	75	73	77	78	78
0.005	73	73	74	78	79	78
0.005	72	74	75	76	77	77
0.005	72	75	77	77	76	79
\bar{X}	72.0	74.2	75.2	76.6	77.4	78.2
0.01	71	74	75	76	77	77
0.01	71	74	75	76	78	77
0.01	72	74	73	78	78	78
0.01	72	75	75	77	76	77
0.01	71	73	76	75	77	75
\bar{X}	71.4	74.0	74.8*	76.4	77.2	76.8*

Table 9a (cont'd).

Relative solubility	Contact time, months					
	3	6	9	12	15	18
0.02	72	74	74	73	76	75
0.02	70	73	74	75	75	76
0.02	71	74	75	76	77	74
0.02	72	73	75	76	77	76
0.02	72	73	73	75	75	74
\bar{X}	71.4	73.4	74.2*	75.0*	76.0*	75.0*

*Significantly different from controls.

Table 9b. Percent weight gain of PVC exposed to aqueous solutions of several organic solutes (second long-term study).

Relative solubility	Contact time, months					
	3	6	9	12	15	18
control	0.17	0.32	0.27	0.08	0.24	0.24
control	0.13	0.25	0.17	0.35	0.25	0.09
control	0.16	0.17	0.09	0.41	0.17	0.13
control	0.25	0.26	0.09	0.34	0.20	0.17
control	0.17	0.34	0.16	0.16	0.17	0.26
\bar{X}	0.18 ^a	0.27 ^a	0.16 ^a	0.27 ^a	0.21 ^a	0.18 ^a
S.D.	0.04	0.07	0.07	0.14	0.04	0.07
0.001	0.17	0.25	0.13	0.26	0.18	0.18
0.001	0.26	0.17	0.17	0.24	0.33	0.27
0.001	0.16	0.33	0.17	0.34	0.25	0.16
0.001	0.17	0.26	0.32	0.24	0.25	0.38
0.001	0.27	0.34	0.17	0.33	0.20	0.07
\bar{X}	0.21 ^a	0.27 ^a	0.19 ^a	0.28 ^a	0.24 ^a	0.21 ^a
S.D.	0.05	0.07	0.07	0.05	0.06	0.12
0.005	0.34	0.33	0.44	0.35	0.41	0.27
0.005	0.34	0.62	0.17	0.17	0.40	0.16
0.005	0.33	0.25	0.25	0.25	0.26	0.18
0.005	0.33	0.25	0.17	0.34	0.34	0.17
0.005	0.45	0.42	0.14	0.17	0.40	0.33
\bar{X}	0.36 ^b	0.37 ^{a,b}	0.23 ^{a,b}	0.26 ^a	0.36 ^b	0.22 ^a
S.D.	0.05	0.15	0.12	0.09	0.06	0.08
0.01	0.33	0.57	0.46	0.41	0.51	0.42
0.01	0.33	0.51	0.31	0.53	0.52	0.42
0.01	0.24	0.33	0.27	0.71	0.59	0.47
0.01	0.28	0.52	0.43	0.62	0.62	0.33
0.01	0.32	0.43	0.20	0.41	0.52	0.21
\bar{X}	0.30 ^b	0.47 ^b	0.33 ^b	0.54 ^b	0.55 ^c	0.37 ^b
S.D.	0.04	0.09	0.11	0.13	0.05	0.10
0.02	0.57	0.64	0.66	0.84	0.73	0.88
0.02	0.57	0.75	0.67	0.82	0.87	0.70
0.02	0.35	0.63	0.68	0.92	0.68	1.01
0.02	0.59	0.58	0.44	0.80	0.74	0.63
0.02	0.54	0.73	0.64	0.91	0.84	0.78
\bar{X}	0.52 ^c	0.67 ^c	0.62 ^c	0.86 ^c	0.77 ^d	0.80 ^c
S.D.	0.10	0.07	0.10	0.05	0.08	0.15

S.D. = standard deviation

Values with different letters, taken at the same time, are significantly different.

Table 10a. Hardness (Barcol) readings of PVC exposed to acetone solutions.

Concentration (%)	Contact time, months					
	3	6	9	12	15	18
control	73	74	80	77	79	81
control	71	74	79	76	77	78
control	72	73	77	77	78	78
control	71	72	78	78	78	79
control	72	74	77	79	78	80
\bar{X}	71.8	73.4	78.2	77.4	78.0	79.2
0.5	71	74	78	77	77	78
0.5	72	75	80	78	78	76
0.5	72	73	77	76	78	79
0.5	71	73	79	77	77	79
0.5	71	72	76	77	79	80
\bar{X}	71.4	73.4	78.0	77.0	77.8	78.4
1.0	71	73	78	77	78	78
1.0	71	74	77	76	79	79
1.0	70	74	78	78	77	79
1.0	71	72	78	76	77	78
1.0	72	73	78	77	78	78
\bar{X}	71.0	73.2	77.8	76.8	77.8	78.4
5.0	71	73	79	76	76	77
5.0	72	72	78	76	76	78
5.0	70	72	79	78	76	77
5.0	71	73	78	77	77	76
5.0	71	74	74	76	75	76
\bar{X}	71.0	72.8	77.6	76.6	76.0*	76.8*
10.0	71	71	76	76	75	75
10.0	71	73	75	74	74	75
10.0	72	72	74	74	75	77
10.0	70	71	76	73	74	76
10.0	71	73	75	75	73	76
\bar{X}	71.0	72.0*	75.2*	74.4*	74.2*	75.8*
20.0	69	71	69	74	72	72
20.0	71	73	73	72	69	71
20.0	70	72	73	73	72	71
20.0	69	71	71	70	69	70
20.0	71	72	72	71	68	70
\bar{X}	70.0*	71.8*	71.6*	72.0*	70.0*	70.8*

*Significantly different from controls.

The results from the first long-term study compare well with these findings. In the previous experiment the Teflon-lined caps were lined with aluminum foil to prevent substantial losses of the organic solutes due to sorption by the Teflon liners. In this experiment there was no aluminum liner. Thus, we concluded that the Teflon does not appear to substantially alter the ability of the test solutions to affect PVC.

Studies on aqueous solutions containing organics that are miscible in water

Long-term acetone study

Table 10 shows the results of long-term (18 months) exposure to 0.5%, 1.0%, 5.0%, 10%, and 20% solutions of acetone. By the end of the study, the 5.0%, 10%, and 20% solutions were able to

Table 10b. Percent weight gain of PVC exposed to acetone solutions.

Concentration (%)	Contact time, months					
	3	6	9	12	15	18
0.0	0.17	0.26	0.28	0.36	0.24	0.17
0.0	0.13	0.18	0.27	0.25	0.26	0.09
0.0	0.17	0.24	0.22	0.17	0.09	0.17
0.0	0.17	0.18	0.26	0.16	0.25	0.08
0.0	0.16	0.08	0.26	0.33	0.35	0.14
\bar{X}	0.16 ^a	0.19 ^a	0.26 ^a	0.25 ^a	0.24 ^a	0.13 ^a
S.D.	0.02	0.07	0.02	0.09	0.09	0.04
0.5	0.17	0.27	0.33	0.31	0.34	0.34
0.5	0.13	0.34	0.25	0.27	0.25	0.22
0.5	0.13	0.24	0.24	0.40	0.26	0.08
0.5	0.17	0.24	0.17	0.28	0.25	0.25
0.5	0.17	0.25	0.27	0.35	0.31	0.33
\bar{X}	0.15 ^a	0.27 ^a	0.25 ^a	0.32 ^{a,b}	0.28 ^a	0.24 ^{a,b}
S.D.	0.02	0.04	0.06	0.06	0.04	0.10
1.0	0.26	0.32	0.33	0.33	0.26	0.27
1.0	0.15	0.25	0.33	0.26	0.27	0.18
1.0	0.27	0.25	0.27	0.26	0.24	0.27
1.0	0.26	0.25	0.25	0.27	0.25	0.20
1.0	0.27	0.25	0.24	0.26	0.25	0.18
\bar{X}	0.24 ^b	0.26 ^a	0.28 ^a	0.28 ^a	0.25 ^a	0.22 ^{a,b}
S.D.	0.05	0.03	0.04	0.03	0.01	0.05
5.0	0.18	0.45	0.42	0.36	0.35	0.33
5.0	0.16	0.26	0.33	0.44	0.32	0.32
5.0	0.25	0.33	0.41	0.40	0.26	0.35
5.0	0.27	0.25	0.27	0.49	0.25	0.34
5.0	0.21	0.17	0.43	0.44	0.34	0.25
\bar{X}	0.21 ^b	0.29 ^a	0.37 ^b	0.43 ^{b,c}	0.30 ^a	0.32 ^b
S.D.	0.04	0.10	0.07	0.05	0.05	0.04
10	0.33	0.50	0.49	0.49	0.51	0.52
10	0.36	0.36	0.66	0.45	0.44	0.45
10	0.26	0.49	0.64	0.70	0.42	0.51
10	0.25	0.60	0.46	0.45	0.50	0.53
10	0.34	0.46	0.40	0.43	0.27	0.40
\bar{X}	0.31 ^c	0.48 ^b	0.53 ^c	0.50 ^c	0.43 ^b	0.48 ^c
S.D.	0.05	0.09	0.11	0.11	0.10	0.06
20	1.08	1.39	1.68	1.95	1.90	1.86
20	1.10	1.45	1.68	1.66	1.95	1.98
20	1.16	1.14	1.69	1.45	1.75	1.95
20	1.10	1.37	1.64	1.81	1.52	2.03
20	1.07	1.25	1.77	1.81	1.86	1.52
\bar{X}	1.10 ^d	1.32 ^c	1.69 ^d	1.74 ^d	1.80 ^c	1.87 ^d
S.D.	0.04	0.12	0.05	0.19	0.17	0.21

S.D. = standard deviation

Values with different letters, taken at the same time, are significantly different.

cause significant changes in the hardness readings and weight gain data.

Study with acetone and methylene chloride

Previously, we (Parker and Ranney 1995) observed that a 40% solution of acetone softened PVC in seven days while a 20% solution did not.

In a different study, we (Parker and Ranney 1994) observed that PVC was not softened after seven days' contact with a solution containing methylene chloride with a relative solubility of 0.4. However, in this study, a solution containing methylene chloride with a relative solubility of 0.4 (0.76%) and 0.5% acetone softened PVC after only four

Table 11a. Hardness (Barcol) readings of PVC exposed to aqueous solutions of methylene chloride (MC) and acetone.

MC, Rel. sol.	Acetone (%)	Contact time, days				
		2	4	7	14	21
control		75	75	74	73	74
control		76	74	76	73	75
control		76	76	73	74	76
control		74	74	75	76	75
control		75	75	76	75	75
\bar{X}		75.2	74.8	74.8	74.2	75
0.4,	0.0	73	72	74	73	74
0.4,	0.0	74	74	74	76	72
0.4,	0.0	76	73	76	73	75
0.4,	0.0	75	75	73	74	72
0.4,	0.0	74	76	73	72	73
\bar{X}		74.4	74	74	73.6	73.2
0.4,	0.5	75	74	71	70	72
0.4,	0.5	76	70	68	71	72
0.4,	0.5	74	67	67	67	70
0.4,	0.5	75	69	70	75	71
0.4,	0.5	75	68	68	70	70
\bar{X}		75	69.6*	68.8*	70.6	71.0*
0.4,	1.0	74	68	66	68	74
0.4,	1.0	74	69	72	72	71
0.4,	1.0	75	71	71	67	71
0.4,	1.0	76	67	69	74	73
0.4,	1.0	75	72	70	73	69
\bar{X}		74.8	69.4*	69.6*	70.8	71.6*
0.4,	5.0	74	71	70	69	72
0.4,	5.0	75	68	70	68	69
0.4,	5.0	74	73	67	71	73
0.4,	5.0	74	68	69	73	72
0.4,	5.0	75	70	70	71	72
\bar{X}		74.4	70.0*	69.2*	70.4*	71.6*
0.4,	10.0	75	72	71	63	69
0.4,	10.0	74	68	69	65	72
0.4,	10.0	74	70	70	69	68
0.4,	10.0	73	66	68	68	67
0.4,	10.0	73	67	70	69	68
\bar{X}		73.8	68.6*	69.6*	66.8*	68.8*
0.4,	20.0	67	63	60	54	53
0.4,	20.0	71	62	54	51	54
0.4,	20.0	68	65	53	52	48
0.4,	20.0	72	58	59	58	46
0.4,	20.0	69	60	67	62	45
\bar{X}		69.4*	61.6*	58.6*	55.4*	49.2*

*Significantly different from controls.

Table 11b. Percent weight gain of PVC exposed to aqueous solutions of methylene chloride (MC) and acetone.

MC, activity	Acetone (%)	Contact time, days				
		2	4	7	14	21
control		0.00	0.16	0.18	0.24	0.16
control		0.00	0.17	0.08	0.33	0.17
control		0.08	0.16	0.17	0.24	0.23
control		0.08	0.25	0.16	0.25	0.16
control		0.00	0.25	0.24	0.32	0.28
\bar{X}		0.03 ^a	0.20 ^a	0.17 ^a	0.28 ^a	0.20 ^a
S.D.		0.04	0.05	0.06	0.04	0.05
0.4,	0.0	0.24	0.32	0.56	0.58	0.71
0.4,	0.0	0.17	0.46	0.55	0.56	0.77
0.4,	0.0	0.26	0.40	0.56	0.61	0.87
0.4,	0.0	0.23	0.55	0.33	0.54	0.76
0.4,	0.0	0.16	0.49	0.56	0.48	0.92
\bar{X}		0.21 ^b	0.44 ^{a,b}	0.51 ^{a,b}	0.55 ^a	0.81 ^b
S.D.		0.04	0.09	0.10	0.05	0.09
0.4,	0.5	0.16	0.57	0.62	0.84	0.70
0.4,	0.5	0.42	0.48	0.76	1.03	0.65
0.4,	0.5	0.16	0.48	0.67	0.80	0.73
0.4,	0.5	0.16	0.58	0.65	0.89	0.82
0.4,	0.5	0.26	0.33	0.63	0.95	1.01
\bar{X}		0.23 ^b	0.49 ^b	0.67 ^b	0.90 ^b	0.78 ^b
S.D.		0.11	0.10	0.05	0.09	0.14
0.4,	1.0	0.25	0.67	0.87	1.20	0.97
0.4,	1.0	0.29	0.83	0.73	1.12	1.01
0.4,	1.0	0.16	0.70	0.51	1.12	1.00
0.4,	1.0	0.47	0.49	0.68	1.07	1.17
0.4,	1.0	0.32	0.64	0.58	0.92	1.20
\bar{X}		0.30 ^{b,c}	0.67 ^b	0.67 ^b	1.09 ^b	1.07 ^{b,c}
S.D.		0.11	0.12	0.14	0.11	0.11
0.4,	5.0	0.42	0.75	0.76	0.91	1.07
0.4,	5.0	0.53	0.70	0.80	0.93	1.27
0.4,	5.0	0.44	0.59	0.80	1.09	1.13
0.4,	5.0	0.33	0.64	0.73	1.22	1.20
0.4,	5.0	0.44	0.73	0.88	1.20	1.08
\bar{X}		0.43 ^c	0.68 ^b	0.79 ^b	1.07 ^b	1.15 ^c
S.D.		0.07	0.07	0.06	0.15	0.08
0.4,	10.0	0.41	1.15	1.81	2.07	2.04
0.4,	10.0	0.50	1.51	1.85	1.65	1.85
0.4,	10.0	0.50	1.12	1.58	1.79	1.87
0.4,	10.0	0.83	0.80	1.51	1.78	1.97
0.4,	10.0	0.82	1.12	1.34	1.65	2.30
\bar{X}		0.61 ^d	1.14 ^c	1.62 ^c	1.79 ^c	2.01 ^d
S.D.		0.20	0.25	0.21	0.17	0.18
0.4,	20.0	2.33	4.53	6.46	6.64	6.34
0.4,	20.0	2.05	5.05	5.51	5.66	6.57
0.4,	20.0	2.00	4.70	5.58	5.91	7.35
0.4,	20.0	2.03	4.38	4.96	6.18	7.71
0.4,	20.0	2.49	4.00	4.70	4.94	7.37
\bar{X}		2.18 ^e	4.53 ^d	5.44 ^d	5.87 ^d	7.07 ^e
S.D.		0.22	0.39	0.68	0.63	0.58

S.D. = standard deviation

Values with different letters, taken at the same time, are significantly different.

days' contact (Table 11). Clearly there is a combined effect with these two organic solutes!

CONCLUSIONS

These data show that when the relative solubility of a single organic solute is less than 0.1, there will not be any measurable effect on PVC, either weight gain or hardness readings. For solutions that contain several organic solvents, we have found the sum of the relative solubilities of the organic solutes in the test solution to be a useful, if not rudimentary, indicator of the ability of an aqueous solution to degrade PVC. If the sum of the relative solubilities of the organic solutes in solution is less than 0.1, there does not appear to be any measurable effect on rigid PVC. This relative solubility corresponds to an activity level (0.1) where diffusion is slowest and Fickian. At slightly higher relative solubilities (0.2, 0.4), there are slight changes in the hardness readings, although in the time frame of our studies (18 months and less) PVC does not become pliable or rubbery.

We do not have any parameters that can be used to predict the ability of aqueous solutions that also contain PVC solvents and swelling agents that are completely miscible in water to soften PVC. The relative solubility cannot be used to predict the ability of the solution to soften PVC because this involves dividing the concentration of the analyte by an infinitely large number. It is interesting that a 1% solution of acetone had no effect on the hardness readings after 18 months' exposure, while a solution containing methylene chloride with a relative solubility of 0.4 (0.76%) and 0.5% acetone softened PVC after only a few days' contact. Clearly, further study is needed to delineate the influence miscible PVC solvents and swelling agents have on the softening ability of aqueous organic solutions.

Previously Leggett and Parker (1994) used a model based on multiparameter linear solvation energy relationships to describe equilibrium partitioning between PVC and water for several organic compounds. Partitioning of these solutes was related to the molecule's acidity, polarity/polarizability, basicity, and molecular volume. A similar approach might prove useful here for predicting in a more meaningful way the ability of aqueous organic solutions to soften PVC.

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APPENDIX A: PERCENT WEIGHT GAIN IN PVC EXPOSED TO VARIOUS NEAT ORGANIC CHEMICALS (HEEREMANS, RANNEY AND PARKER; UNPUBLISHED STUDIES)*

<i>Solvent</i>	<i>Percent Δ wt.</i>			
	<i>4 hours</i>	<i>1 day</i>	<i>7 days</i>	<i>55 days</i>
acetic acid	0.0	0.2	—	—
acetone	46.5	—	—	150.8
acetonitrile	1.8	5.0	11.2	13.9
benzene	2.2	5.8	23.3	47.5
benzoinitrile	26.8	—	—	365.0
benzyl alcohol	0.0	0.0	0.0	0.1
bromochloromethane	309.3	648.4	—	—
bromodichloromethane	60.8	216.2	—	—
bromoform	20.8	71.1	—	—
butanol	0.0	0.0	0.0	0.0
butylamine	14.5	59.3	—	—
carbon disulfide	12.0	28.0	36.5	33.9
carbon tetrachloride	0.0	0.0	0.0	0.1
chlorobenzene	11.7	52.1	—	149.7
chlorodibromomethane	57.2	165.3	—	—
2-chloroethyl vinyl ether	7.5	25.6	—	—
chloroform	36.2	124.2	—	205.1
cyclohexane	0.0	0.1	0.0	0.0
cyclohexanone	D			
1,2-dibromoethane	27.8	105.4	—	—
1,2-dichlorobenzene	3.6	17.7	61.1	197.4
1,3-dichlorobenzene	3.6	13.9	47.5	100.8
1,1-dichloroethane	13.4	44.1	115.1	115.2
1,2-dichloroethane	52.6	215.5	—	—
<i>cis</i> -1,2-dichloroethylene	87.3	253.0	—	—
<i>trans</i> -1,2-dichloroethylene	18.7	50.7	—	—
1,2-dichloropropane	10.5	30.8	—	—
3,4-dimethyl acetophenone	—	238.9	—	—
dimethyl formamide	32.5	D		
dimethyl sulfoxide	4.7	17.0	55.4	68.9
dioxane	4.2	14.3	51.0	104.7
ethanol amine	0.0	0.4	—	—
ethyl acetate	23.6	85.4	—	144.4
ethyl benzene	0.8	2.1	—	—
heptane	0.0	0.0	0.0	0.0
hexane	0.0	0.0	0.0	0.0
iso-butanol	0.0	-0.1	-0.1	0.1
iso-propanol	0.0	0.0	0.0	-0.1
methanol	0.0	0.0	0.1	0.3
methoxy ethanol	0.1	0.9	—	—
methyl ethyl ketone	—	—	—	238.8

* Small pieces of PVC well casing (1 cm \times 1 cm) were placed in 22-mL borosilicate glass vials, 5 mL of solvent was added, and the vials were capped with Teflon-lined plastic caps. Prior to weighing (\pm 0.0001), the samples were blotted with paper towels.

— No data

D PVC dissolved

<i>Solvent</i>	<i>Percent Δ wt.</i>			
	<i>4 hours</i>	<i>1 day</i>	<i>7 days</i>	<i>55 days</i>
methyl isobutyl ketone	18.5	92.3	—	—
methylene chloride	—	—	—	388.0
nitrobenzene	29.6	—	—	488.3
<i>o</i> -nitrotoluene	11.9	46.1	—	474.5
<i>m</i> -nitrotoluene	14.5	51.1	—	479.6
propanol	0.0	0.0	0.0	0.0
pyridine	122.8	D		
1,1,2,2-tetrachloroethane	6.1	32.3	—	—
tetrachloroethylene	0.1	0.1	0.3	0.3
tetrahydrofuran	22.9	D		
toluene	2.7	8.0	32.0	49.2
1,1,1-trichloroethane	0.2	0.5	1.6	4.0
1,1,2-trichloroethane	22.8	83.0	328.8	368.5
trichloroethylene	7.6	27.8	69.7	67.8
<i>o</i> -xylene	0.7	2.4	—	—
<i>m</i> -xylene	0.6	2.3	—	—
<i>p</i> -xylene	1.2	2.7	—	—

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